

The Building Blocks of Battery Technology:

Using modified tower block game sets to explain and aid the understanding of rechargeable Li-ion batteries

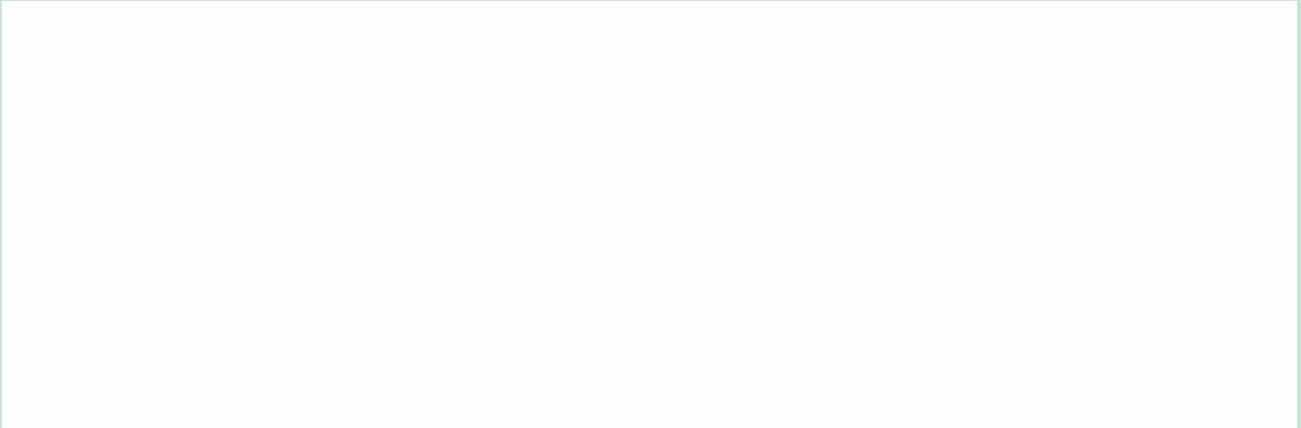
Activity Sheet: GCSE Chemistry and Scottish National Qualifications

Introduction

Lithium ion (Li-ion) batteries are a rechargeable-type of battery which have become a staple in modern-day life and are used in mobile phones, laptops and, in more recent times, electric vehicles. As these batteries are rechargeable they have to be charged, used (discharged) and then recharged. The majority of electrode materials currently used in Li-ion batteries have a layered structure, and so Jenga can be used to help demonstrate charging and discharging processes, in addition to exploring why batteries fail over time and why the rate of charge is important. The original Li-ion batteries used LiCoO_2 and graphite as the electrodes. In newer cells, some of the cobalt is typically replaced by nickel and manganese in order to reduce the cost (cobalt is a less abundant element).

Based on the introduction paragraph above, what one word describes the type of battery Li-ion batteries are.

Graphite is used as an electrode in a Li-ion battery. Draw the structure of graphite, either from your notes or from a Google search. What are the characteristic properties of graphite? Advanced level - from the structure you have drawn, where do you think Li-ions could go in the structure?



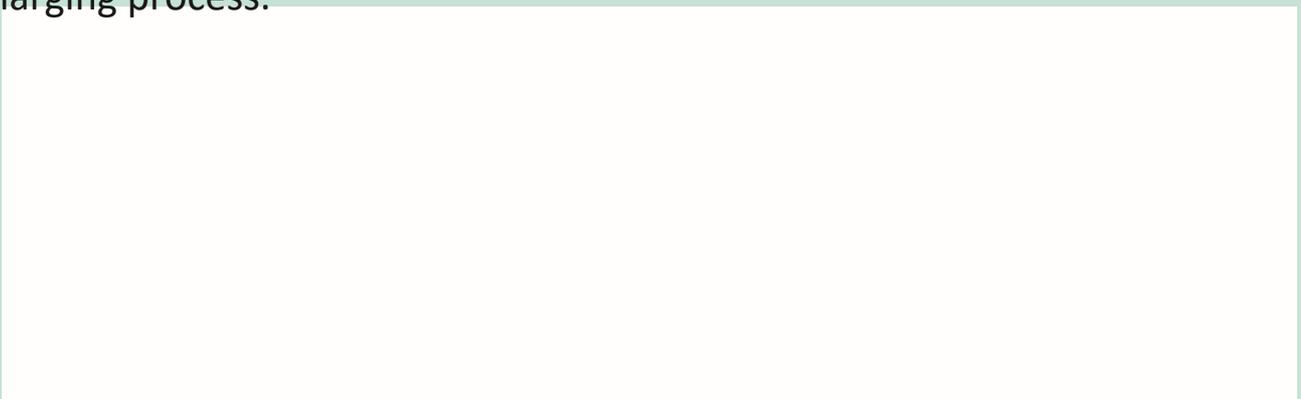
What to investigate:

1) Charging-Discharging

When batteries are charging, Li-ions move from the oxide electrode to the graphite electrode. The reverse process occurs on discharge.

- Have a go at charging your battery Jenga by removing the Li-ion blocks from the oxide electrode and inserting them between the layers of the graphite sheets on the graphite electrode (i.e. in place of the white blank blocks).
- Do the reverse process to reset the battery Jenga to show discharge.

Draw a diagram of the two electrodes you have set up. Then draw arrows of where you have moved the Li-ions from one electrode to the other on a charging process.



2) Rate of Charge

Think about how long mobile phones take to charge - do you think electric cars take the same time to charge?

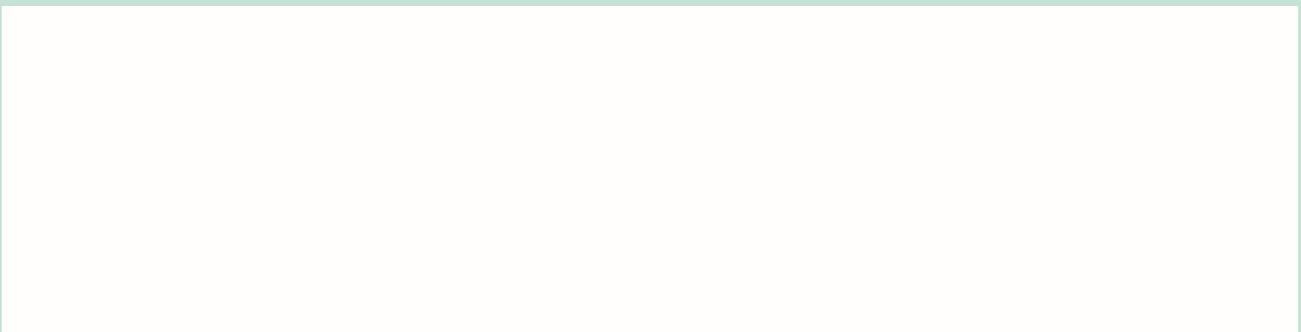
- Focusing on the oxide electrode, time yourself removing a Li-ion block every 10 seconds.
- How many Li-ions are you able to remove in the space of 30 seconds?
- How many could you remove in total, without collapsing the structure?

Place all the Li-ion blocks back into the oxide electrode, so we have reset the structure.

Again, focusing on the oxide electrode, how many Li-ions can you remove in 30 seconds, if you remove the blocks every couple of seconds?

(Don't worry if the structure collapses on this fast charge).

In your own words, describe what you have observed from this demonstration and why it is important to have a high charge rate for applications, but how this can impact safety.



3) Degradation

In the next demonstration we're going to think about the lifetime of the batteries.

- In the first demonstration where we showed charging and discharging (known as cycling our battery), did you find it tricky to remove the blocks?
- Did any of the electrodes collapse or were knocked quite easily?

Cycle (charge and discharge) your battery Jenga set multiple times.

- Can it help you work out why batteries stop working as well over time?

Draw a diagram of how the oxide electrode looks after you've repeatedly cycled the battery Jenga.



Write 5 take-home points for the whole of this activity.

Li-ion batteries store electrochemical energy. What forms of energy generation could we use to charge our battery? State several different energy generation sources. If you have time, think about the energy form, i.e wind power - wind turbines convert kinetic energy to electrical energy.



What physical properties do you think Li-ion batteries require? Think about mobile phones - are they heavy or light?

